

$$\boxed{1} \quad C = \frac{m_s (g)}{V_D (l)} = \frac{124 g}{2 l} = \underline{\underline{62 \text{ g/l}}}$$

$$\boxed{2} \quad m_s = 0,15 \text{ kg} = 150 \text{ g}$$

$$V_D = 5 \text{ dm}^3 = 5 \text{ l}$$

$$C = \frac{m_s (g)}{V_D (l)} = \frac{150}{5} = \underline{\underline{30 \text{ g/l}}}$$

$$\boxed{3} \quad C = \frac{m_s \cdot 100}{m_D} = \frac{125 \cdot 100}{125 + 2500} = \underline{\underline{4,76 \%}}$$

$$\boxed{4} \quad C = \frac{m_s \cdot 100}{m_D} = \frac{50 \cdot 100}{500} = \underline{\underline{10 \%}}$$

$$m_s = 50000 \text{ mg} = 50 \text{ g}$$

$$m_{\text{disolvente}} = 0,45 \text{ kg} = 450 \text{ g}$$

$$m_D = 50 + 450 = 500 \text{ g} \rightarrow \text{masa de la } \underline{\underline{\text{disolución}}}$$

$$\boxed{5} \quad C = \frac{V_s \cdot 100}{V_D} = \frac{200 \cdot 100}{200 + 750} = \underline{\underline{21 \%}}$$

(No es necesario pasar V_s o V_D a litros,
 ya que usamos la misma unidad (ml)
 en el numerador y en el denominador)

$$\boxed{6} \quad V_S = 125 \text{ ml} \quad V_{\text{disolvente}} = 250 \text{ ml}$$

$$V_D = 250 \text{ ml} + 125 \text{ ml} = 375 \text{ ml}$$

$$C = \frac{V_S \cdot 100}{V_D} = \frac{125 \cdot 100}{375} = \boxed{33,3\%}$$

$$\boxed{7} \quad m_S = 250 \text{ g}$$

$$V_D = 10000 \text{ cm}^3 = 10 \text{ dm}^3 = 10 \text{ l}$$

$$C = \frac{m_S}{V_D} = \frac{250}{10} = \boxed{25 \text{ g/l}}$$

$$\boxed{8} \quad V_S = 250 \text{ ml} = 0,25 \text{ l}$$

$$V_D = 2 \text{ l}$$

} Tienen que estar
ambos volúmenes en
la misma unidad.

$$C = \frac{V_S \cdot 100}{V_D} = \frac{0,25 \cdot 100}{2} = \boxed{12,5\%}$$

$$\boxed{9} \quad m_S = 120 \text{ g}$$

$$m_D = 0,75 \text{ kg} = 750 \text{ g}$$

} Ambas masas en
la misma
unidad.

$$C = \frac{m_S \cdot 100}{m_D} = \frac{120 \cdot 100}{750} = \boxed{16\%}$$